

(19) Europäisches Patentamt
European Patent Office
Office européen des brevets



(11) EP 0 659 478 B1

(12)

EUROPEAN PATENT SPECIFICATION

(45) Date of publication and mention
of the grant of the patent:
18.03.1998 Bulletin 1998/12

(51) Int. Cl.⁶: B01J 21/12

(21) Application number: 94203703.7

(22) Date of filing: 20.12.1994

(54) Process for preparing amorphous, catalytically active silico-aluminas

Verfahren zur Herstellung von amorphen, katalytisch aktiven Silikoaluminiumoxyden

Procédé de préparation de silico-alumines catalytiquement actives amorphes

(84) Designated Contracting States:
AT BE CH DE DK ES FR GB GR IE LI LU MC NL
PT SE

Designated Extension States:
SI

(30) Priority: 22.12.1993 IT MI932696

(43) Date of publication of application:
28.06.1995 Bulletin 1995/26

(73) Proprietors:

- ENIRICERCHE S.p.A.
20097 S. Donato Milanese (Milano) (IT)
- AGIP PETROLI S.p.A.
00142 Roma (IT)
- ENICHEM S.p.A.
20124 Milano (IT)

(72) Inventors:

- Perego, Carlo
I-20040 Carnate (Milano) (IT)
- Peratello, Stefano
I-20054 Nova Milanese (Milano) (IT)
- Millini, Roberto
I-20077 Cerro al Lambro (Milano) (IT)

(74) Representative:

Fusina, Gerolamo et al
Ing. Barzanò & Zanardo Milano S.p.A.,
Via Borgonuovo, 10
20121 Milano (IT)

(56) References cited:

EP-A- 0 340 868 EP-A- 0 408 300

- JOURNAL OF NON-CRYSTALLINE SOLIDS, vol.
147/148, 01 October 1992; LOPEZ et al., pp. 769-
772

EP 0 659 478 B1

Note: Within nine months from the publication of the mention of the grant of the European patent, any person may give notice to the European Patent Office of opposition to the European patent granted. Notice of opposition shall be filed in a written reasoned statement. It shall not be deemed to have been filed until the opposition fee has been paid. (Art. 99(1) European Patent Convention).

Description

The present invention relates to a process for preparing an amorphous and micro/mesoporous silica-alumina gel having a large surface area and a controlled pore size, which is catalytically active in hydrocarbon conversion reactions.

Some silica and alumina gels displaying catalytic activity are known in the art. So, e.g., EP-A-160 145 discloses a process for alkylating aromatic hydrocarbons using, as the catalyst, a silica/alumina gel of amorphous character with a diameter of the pores typically of 50-500 Ångstroms and with a molar silica:alumina ratio typically comprised within the range of from 1:1 to 10:1.

Furthermore, R.M.S. Malton and J. Davidz in *Journal of Catalysis*, 60, 156-166 (1979) disclose a process for the synthesis of amorphous silica and alumina catalysts, with a controlled pore volume.

Finally, US-A-5,049,536 discloses a microporous, x-ray amorphous silica/alumina gel having a surface area of from 500 to 1000 m²/g and a pore volume comprised within the range of from 0.3 to 0.6 cm³/g. Such a silica and alumina gel, which is catalytically active in hydrocarbon preparation reactions, is typically prepared by hydrolysing a tetra-alkyl ammonium hydroxide, a hydrolysable aluminum compound and a hydrolysable silicon compound and causing the resulting hydrolysis mixture to undergo gelation by operating in an aqueous media and at low temperatures, and then submitting the resulting gel to drying and firing.

The present Applicant found now, according to the present invention, that, in relation with US-A-5,049,036, special conditions exist for the hydrolysis of the above mentioned reactants and subsequent gelation which make it possible a silica-alumina gel to be obtained which is unexpectedly improved in terms of control of its surface characteristics, in particular its porous structure, as well as in terms of its catalytic performance, in particular its activity and useful life time under use conditions.

In accordance therewith, the present invention relates to a process for preparing an amorphous, micro/mesoporous silica-alumina gel with controlled pore size, having a surface area of at least 500 m²/g and with a molar ratio of SiO₂:Al₂O₃ of at least 30:1, by starting from a tetra-alkyl ammonium hydroxide, an aluminum compound capable of yielding Al₂O₃ by hydrolysis (i.e., hydrolysable to Al₂O₃), and a silicon compound capable of yielding SiO₂ by hydrolysis (i.e., hydrolysable to SiO₂), whereby:

-- said tetra-alkyl ammonium hydroxide is a tetra-(C₂-C₅)-alkyl ammonium hydroxide, said hydrolysable aluminum compound is an aluminum tri-(C₂-C₄)-alkoxide and said hydrolysable silicon compound is a tetra-(C₁-C₅)-alkyl orthosilicate, and said reactants are submitted to hydrolysis and gelation, by operating at a temperature equal to, or higher than, the boiling temperature, under atmos-

pheric pressure, of any alcohol which are developed as a by-product from said hydrolysis reaction, without said alcohols being removed, or being substantially removed, from the reaction media; and -- the so produced gel is dried and fired.

The useful tetra-alkyl ammonium hydroxide for the purposes of the present invention is therefore selected from tetraethyl-, propyl-, isopropyl-, butyl-, isobutyl-, t-butyl-, and pentyl-ammonium hydroxide, and among these tetra-propyl-, tetraisopropyl- or tetrabutyl-ammonium hydroxide are preferred.

Aluminum trialkoxide is selected from aluminum triethoxide, propoxide, isopropoxide, butoxide, isobutoxide and t-butoxide. Among these, aluminum tripropoxide and tri-isopropoxide are preferred.

Tetra-alkyl orthosilicate is selected from tetramethyl-, tetraethyl-, propyl-, isopropyl-, butyl-, isobutyl-, t-butyl- and pentyl-orthosilicate and among these, tetraethyl orthosilicate is preferred.

The process according to the present invention is carried out by first preparing an aqueous solution containing tetra-alkyl ammonium hydroxide and aluminum tri-alkoxide, by operating at a high enough temperature is order to secure that the aluminum compound will be properly dissolved. To said aqueous solution, tetra-alkyl orthosilicate is then added. The resulting mixture is then heated up to a suitable temperature for starting the hydrolysis reaction. Said temperature is a function of the composition of the reaction mixture (comprised within the range of from 70 to 100°C). The hydrolysis reaction is an exothermic one and therefore once that the reaction is started, its self-feeding is secured. The amounts of the reactants which constitute the reaction mixture should furthermore be such as to comply with the molar ratios: SiO₂:Al₂O₃ comprised within the range of from 30:1 to 500:1, tetra-alkyl ammonium hydroxide:SiO₂ comprised within the range of from 0.05:1 to 0.2:1 and H₂O:SiO₂ comprised within the range of from 5:1 to 40:1. Preferred values for these molar ratios are: SiO₂:Al₂O₃ from 50:1 to 300:1; tetra-alkyl ammonium hydroxide:SiO₂ from 0.05:1 to 0.2:1 and H₂O:SiO₂ from 10:1 to 25:1.

As said above, the basic aspect of the process of the present invention is that the hydrolysis of the reactants and their gelation is carried out by operating at a temperature equal to, or higher than, the boiling temperature, under atmospheric pressure, of any alcohols which are developed as a by-product from said hydrolysis reaction, without said alcohols being removed, or substantially removed, from the reaction media. Therefore, the hydrolysis/gelation temperature results to be critical and will be properly kept comprised within the range of from about >65°C up to about 110°C. Furthermore, in order to keep any developed alcohols retained in the reaction media, the process can be carried out in an autoclave under the autogenous system pressure at the selected operating temperature (normally of the

order of 1.1-1.5 abs.bars), or the process can be carried out under the atmospheric pressure inside a reactor equipped with a reflux condenser.

According to a particular embodiment of the process, the hydrolysis and gelation are carried out in the presence of a larger alcohol amount than as developed as a reaction by-product. For this purpose, a free alcohol, and preferably ethanol, is added to the reaction mixture up to a maximal value of molar ratio of added alcohol:SiO₂, of 8:1.

The necessary time in order to cause the hydrolysis and gelation to reach their completion, under the above shown conditions, is normally comprised within the range of from 10 minutes to 3 hours, and preferably is of the order of 1-2 hours.

It was furthermore found useful to submit the resulting gel to ageing, by keeping the hydrolysis/gelation mixture in the presence of the alcohol and under room temperature conditions for a time of the order of from 1 to 24 hours.

Finally, the alcohol is removed from the gel, and the latter is dried under vacuum (e.g., under a vacuum of 30 torr), at a temperature of 110°C. The dried gel is finally fired under an oxidizing atmosphere (normally atmospheric air) at a temperature comprised within the range of from 500 to 700°C during a time period of from 4 to 20 hours, and preferably at 500-600°C during 6-10 hours.

The so obtained silica-alumina gel shows a composition which corresponds to the composition of the starting reactants, considering that the reaction yields are practically quantitative. Therefore, the molar ratio of SiO₂:Al₂O₃ will be comprised within the range of from 30:1 to 500:1 and preferably of from 50:1 to 300:1, with most preferred values being of the order of 100:1.

When it is analysed by powder x-ray diffractometry, this gel results to be amorphous; it displays a surface area of at least 500 m²/g and normally comprised within the range of from 600 to 850 m²/g and a pore volume of 0.4-0.8 cm³/g. Finally, it was found that by means of the process of the present invention, the size of the pores can be controlled to be within a certain range of values, and in particular within the range of from 10 to about 20 $\times 10^{-10}$ m (10 to about 20 Ångstroms) (values referred to pore radius), with, in particular, said pores showing a narrow size distribution, as it will be clearer from the experimental examples reported in the following.

The gel obtained according to the present process is an active catalyst in the usual reactions of conversion of hydrocarbons, like light olefin isomerization and oligomerization reactions.

The gel is particularly useful in propylene oligomerization, to yield hydrocarbon cuts, liquid under room conditions, which are constituted by branched olefin oligomers useful for gasoline or jet fuel formulation.

The following Examples are reported in order to better illustrate the invention.

Example 1

In this Example, the process according to the present invention is carried out with the following molar ratios of the components of the starting reaction mixture:

10	- SiO ₂ :Al ₂ O ₃	= 100
	-- TPA-OH:SiO ₂	= 0.09
	-- H ₂ O:SiO ₂	= 15

15 An amount of 4,727 g of water and 3,922 g of TPA-OH (tetra-propyl ammonium hydroxide; used as a solution at 14% by weight in water) are charged to an autoclave of 20 litres of capacity.

20 The solution in the autoclave is heated and when the temperature reaches 60°C, 120 g of Al(OiPr)₃ (aluminum isopropoxide; supplied by Fluka) is added.

25 The autoclave is closed, stirring is started, and the reaction mixture is kept at the above temperature during the necessary time for aluminum compound to be completely dissolved (about 40 minutes). Now, the temperature of the autoclave is increased up to 90°C, and 6,250 g of TEOS (tetra-ethyl orthosilicate) is added through a purposely provided valve. When addition is complete, the valve is closed, the temperature is adjusted at 88°C 30 and the autoclave is kept under said conditions during 3 hours. The pressure reading on the pressure gauge reaches the maximal value of 1.5 bars.

35 A thick liquid product is thusly obtained which, after being aged during approximately 12 hours, is dried inside a rotary dryer, under vacuum (about 60 torr) and is then calcined in air at 550°C for about 8 hours.

40 The analysis by powder x-ray diffractometry [carried out by means of a Philips Vertical Diffractometer equipped with a proportional pulse counter and operating with Cu K- α radiation ($\lambda = 1.54178 \text{ \AA}$)], indicates that the calcined solid product is amorphous.

45 The specific surface area of said solid product [as determined by BET analysis by N₂ adsorption at liquid N₂ temperature (77°K) with a Carlo Erba Sorptomatic 1900 apparatus] resulted to be of 656 m²/g.

50 The specific pore volume [as determined by N₂ adsorption/desorption cycles at 77°K, using a Carlo Erba Sorptomatic 1900 apparatus] is of 0.473 cm³/g, with the pore size distribution being as shown in Figure 1.

55 By adopting the terms suggested by IUPAC Manual of Symbols and Terminology, Appendix 2, Part I Coll., Surface Chem. Pure Appl. Chem., 31, 578 (1972), according to which those pores having a diameter of $<20 \times 10^{-10}$ m (Å) are defined as "micropores" and those pores having a diameter comprised within the range of from 20×10^{-10} m (Å) to 500×10^{-10} m (Å) are defined as "mesopores", the solid product obtained according to

the present invention can be defined as being a "micro/mesoporous" material.

In following Table 1, the data are reported of catalytic activity in propylene oligomerization reaction, operating with a WHSV (weight hourly space velocity) value of 4 h^{-1} , and under a 35-bar pressure.

In this Table, as well as in following tables, the term "T.O.S." means "Time On Stream", i.e., the cumulated run time expressed as hours, and "Conversion (%)" is the calculated conversion rate based on fed propylene.

Table 1

T.O.S.	Temperature (°C)	Conversion (%)
42	140	65
63	150	67
134	190	67

Example 2

The process is carried out by operating according to Example 1, in a 5-litre autoclave, to which the following amounts of reactants are charged:

-- H ₂ O	= 1,182 g
-- TPA-OH	= 980 g
-- Al(OiPr) ₃	= 30 g
-- TEOS	= 1,560 g

The end product obtained from this test run results to be amorphous when analysed by powder x-ray diffractometry; it has a BET surface area of $710 \text{ m}^2/\text{g}$, and a pore volume of $0.596 \text{ cm}^3/\text{g}$.

In following Table 2, the data are reported of catalytic activity in propylene oligomerization reaction, operating with a WHSV value of 4 h^{-1} , and under a 35-bar pressure.

Table 2

T.O.S.	Temperature (°C)	Conversion (%)
20	120	77
63	140	77
133	160	57

Example 3

In this Example, the process is carried out in a 1-litre autoclave, with the following molar ratios of the components of the reaction mixture:

-- SiO ₂ :Al ₂ O ₃	= 100
-- TPA-OH:SiO ₂	= 0.09
-- H ₂ O:SiO ₂	= 15

From 158 g of water, 131 g of TPA-OH (used an a solution at 14% by weight in water) and 4 g of Al(OiPr)₃ a solution is prepared on a heater plate at 60°C . When the aluminum salt is dissolved, the solution is charged to a 1-litre autoclave previously thermostatted at 60°C , through a pin valve.

Now, the temperature of the solution is increased up to about 85°C , and, still through the pin valve, 208 g of TEOS is added.

When the hydrolysis reaction is complete, the reaction mixture is kept at $82-83^\circ\text{C}$ during 8 hours. The pressure reading on the pressure gauge reaches a peak value of 1.4 bars.

A thick liquid product is thus obtained which, after an approximately 12-hour ageing, is dried inside a rotary dryer, under vacuum (about 60 torr) and is then calcined in air at 550°C for about 8 hours.

When analysed by powder x-ray diffractometry, the calcined solid product results to be amorphous; it displays a BET specific surface area of $682 \text{ m}^2/\text{g}$, and a pore volume of $0.537 \text{ cm}^3/\text{g}$.

In following Table 3, the data are reported of catalytic activity in propylene oligomerization reaction, carried out by operating with a WHSV value of 4 h^{-1} , and under a 35-bar pressure.

Table 3

T.O.S.	Temperature (°C)	Conversion (%)
20	150	87
115	150	67
158	150	60

Example 4

In this Example, the process according to the present invention is carried out with the following molar ratios of the components of the reaction mixture:

-- SiO ₂ :Al ₂ O ₃	= 100
-- TPA-OH:SiO ₂	= 0.09
-- H ₂ O:SiO ₂	= 15

An amount of 302 g of water and 274 g of TPA-OH (used an a solution at 14% by weight in water) are

charged to a flask of 2 litres of capacity, equipped with a reflux condenser.

When temperature reaches the value of 50-60°C, 8 g of $\text{Al}(\text{O}i\text{Pr})_3$ is added. When aluminum salt is dissolved, the temperature is increased up to 98°C, heating is discontinued and 416 g of TEOS is added. When the hydrolysis reaction is complete, temperature begins to spontaneously decrease: now, heating is started up again, so as to keep the reaction mixture at the temperature of 82-83°C during 1 hour and 45 minutes.

After being submitted to a 20-hour ageing, the product is discharged from the flask, is dried inside a rotary dryer and in a vacuum oven and is then calcined in air at 550°C for about 8 hours.

When analysed by powder x-ray diffractometry, the calcined solid product results to be amorphous; it displays a BET specific surface area of 804 m^2/g , and a specific pore volume of 0.773 cm^3/g . The result from the porosimetric analysis is reported in Figure 2.

In following Table 4 the data are reported of catalytic activity in propylene oligomerization reaction, carried out by operating with a WHSV value of 4 h^{-1} , and under a 35-bar pressure.

Table 4

T.O.S.	Temperature (°C)	Conversion (%)
25	150	88
54	150	82
134	150	78

Example 5

In this Example, the process is carried out with the following molar ratios of the components of the reaction mixture:

-- $\text{SiO}_2:\text{Al}_2\text{O}_3$	= 100
-- TBA-OH:SiO_2	= 0.09
-- $\text{H}_2\text{O:SiO}_2$	= 15

** TBA = Tetrabutyl ammonium hydroxide, supplied as a 18.9% solution, by weight, in water.

The process is carried out in a reactor equipped with reflux condenser, by operating according to the same modalities as of Example 4, with the following amounts of reactants:

-- water	= 186.5 g
-- TBA-OH	= 103 g
-- $\text{Al}(\text{O}i\text{Pr})_3$	= 4 g
-- TEOS	= 208 g

10 The resulting solid product is amorphous on x-ray analysis; it displays a BET surface area of 837 m^2/g and a pore volume of 0.737 cm^3/g .

Example 6

15 The process is carried out in a reactor equipped with reflux condenser, similarly to preceding Example 4, with the difference that ethyl alcohol (EtOH) is added, previously dissolved in TEOS is added with a molar ratio of:

-- EtOH:TEOS	= 4.
--------------	------

25 The solid product results to be amorphous on x-ray analysis, and shows a BET surface area of 674 m^2/g and a pore volume of 0.552 cm^3/g .

30 In accompanying Figure 4 the microphotograph obtained by transmission electron microscope is reported. A regular distribution and a pore uniformity in sample particles are evidenced. Such an observation is consistent with the porosimetric analysis reported in accompanying Figure 3.

35 From the charts reported in Figures 1-3, one will observe that the samples prepared according to the process of the present Application display an extremely narrow and regular distribution of pore size, with pore radius being of about 20 Ångstrom. The presence of 40 such pores is clearly visible when the product is viewed under transmission electron microscope (TEM). These pores result to be regular as regards their size, and distributed throughout amorphous silicoalumina particles.

45 Example 7 (Comparison Example)

The test is run according to such modalities as disclosed in US 5,049,536; by charging the reactant to a beaker, in the following molar proportions:

-- $\text{SiO}_2:\text{Al}_2\text{O}_3$	= 50
-- TPA-OH:SiO_2	= 0.09
-- $\text{H}_2\text{O:SiO}_2$	= 15

50 ** TPA = tetrapropyl ammonium hydroxide, supplied as a solution at 13.35% by weight in water.

An x-ray analysis amorphous solid product is obtained, which displays a BET surface area of 672 m²/g and a pore volume of 0.454 cm³/g.

In following Table 5, the data are reported of catalytic activity displayed by the product in propylene oligomerization reaction carried out by operating with a WHSV value of 4 h⁻¹, and under a 35-bar pressure.

Table 5

T.O.S.	Temperature (°C)	Conversion (%)
41	120	59
162	150	50
210	170	49

Example 8

In accompanying Figure 5 two life tests are compared, which were carried out: the first one with the sample from Example 4, the other one with the sample from Comparison Example 7.

The characteristics and operating modalities in the test runs were as follows:

- catalyst shape: pelletized catalyst;
- catalyst size: 841 x m - 370 x m (20-40 mesh);
- reactor type: fixed bed reactor;
- feed: propylene/propane (70:30 by weight) mixture;
- reactor temperature: from 120°C to 180°C;
- reactor pressure: 38 bars;
- weight hourly space velocity WHSV: 4 g of propylene per active phase gram per hour;
- total reaction time: about 240 hours.

The tests were carried out by stepwise increasing the reaction temperature, by 10°C each time (about every 24 hours) in order to compensate for the conversion decrease due to catalyst deactivation.

As one will see from Figure 5, silicoalumina from Example 4 ("--o--" line) constitutes a catalyst with improved performance over silicoalumina from Comparison Example 7 ("---*---" line).

In particular, the value of productivity (understood as grams of oligomer produced per each gram of active phase) of the catalyst from Example 4 is of 510 g/g, whilst for catalyst from Comparison Example 7, it is of 400 g/g.

Claims

1. Process for preparing an amorphous, micro/mesoporous silica-alumina gel with controlled pore size, having a surface area of at least 500 m²/g and with a molar ratio of SiO₂:Al₂O₃ of at least 30:1, by starting from a tetra-alkyl ammonium hydroxide, an alu-

minum compound capable of yielding Al₂O₃ by hydrolysis (i.e., hydrolysable to Al₂O₃), and a silicon compound capable of yielding SiO₂ by hydrolysis (i.e., hydrolysable to SiO₂), WHEREBY:

-- said tetra-alkyl ammonium hydroxide is a tetra-(C₂-C₅)-alkyl ammonium hydroxide, said hydrolysable aluminum compound is an aluminum tri-(C₂-C₄)-alkoxide and said hydrolysable silicon compound is a tetra-(C₁-C₅)-alkyl orthosilicate, and said reactants are submitted to hydrolysis and gelation, by operating at a temperature equal to, or higher than, the boiling temperature, under atmospheric pressure, of any alcohols which are developed as a by-product from said hydrolysis reaction, without said alcohols being removed, or being substantially removed, from the reaction media; and
-- the so produced gel is dried and fired.

2. Process according to claim 1, characterized in that said tetra-alkyl ammonium hydroxide is tetra-propyl-, tetraisopropyl- or tetrabutyl-ammonium hydroxide, said aluminum trialkoxide is aluminum tri-propoxide and tri-isopropoxide and said tetra-alkyl orthosilicate is tetra-ethyl orthosilicate.
3. Process according to claim 1 or 2, characterized in that an aqueous solution is prepared which contains tetra alkyl ammonium hydroxide and aluminum trialkoxide and to said aqueous solution tetra alkyl orthosilicate is added, by operating at a lower temperature than the hydrolysis temperature, and with such amounts of the starting reactants as to comply with the molar ratios: SiO₂:Al₂O₃ comprised within the range of from 30:1 to 500:1, tetra alkyl ammonium hydroxide:SiO₂ comprised within the range of from 0.05:1 to 0.2:1 and H₂O:SiO₂ comprised within the range of from 5:1 to 40:1, and hydrolysis and gelation is caused by heating up to a temperature comprised within the range of from about >65°C up to about 110°C, in an autoclave under the autogenous system pressure, or under atmospheric pressure inside a reactor equipped with a reflux condenser.
4. Process according to claims 1-3, characterized in that a free alcohol, and preferably ethanol, is added to the reaction mixture up to a maximal value of molar ratio of added alcohol:SiO₂, of 8:1.
5. Process according to claims 1-4, characterized in that the hydrolysis/gelation time is comprised within the range of from 10 minutes to 3 hours, and preferably is of the order of 1-2 hours.
6. Process according to claims 1-5, characterized in that the gel is submitted to ageing by keeping the

hydrolysis/gelation mixture in the presence of the alcohol and under room temperature conditions for a time of the order of from 1 to 24 hours.

7. Process according to any of preceding claims, characterized in that the alcohol is removed from the gel, and the latter is dried by operating under vacuum and the dried gel is then fired under an oxidizing atmosphere at a temperature comprised within the range of from 500 to 700°C during a time period of from 4 to 20 hours, and preferably at 500-600°C during 6-10 hours. 10

8. Use of silica-alumina gel as obtained according to claims 1 to 7, in hydrocarbon conversion processes. 15

9. Use according to claim 8, in which said hydrocarbon conversion process is propylene oligomerization. 20

Patentansprüche

1. Verfahren zur Herstellung eines amorphen mikro-/mesoporösen Siliciumdioxid-Aluminiumoxid-Gels mit geregelter Porengröße, das eine Oberfläche von mindestens 500 m²/g und ein Molverhältnis von SiO₂:Al₂O₃ von mindestens 30:1 hat, durch Verwendung eines Tetraalkylammoniumhydroxids, einer Aluminiumverbindung, die zur Bildung von Al₂O₃ durch Hydrolyse befähigt ist (d.h. zu Al₂O₃ hydrolysierbar ist), und einer zur Bildung von SiO₂ durch Hydrolyse befähigten Siliciumverbindung (d.h. die zu SiO₂ hydrolysierbar ist) als Ausgangsmaterialien, wobei 25

das Tetraalkylammoniumhydroxid ein Tetra-(C₂-C₅)-alkylammoniumhydroxid ist, die hydrolysierbare Aluminiumverbindung ein Aluminium-tri-(C₂-C₄)-alkoxid ist und die hydrolysierbare Siliciumverbindung ein Tetra-(C₁-C₅)-alkylorthosilicat ist, und die Reaktanten der Hydrolyse und Gelbildung unterworfen werden, indem bei einer Temperatur gearbeitet wird, die gleich der Siedetemperatur unter atmosphärischem Druck der Alkohole, die als Nebenprodukt durch die Hydrolysereaktion gebildet werden, oder die höher als diese Temperatur ist, ohne daß diese Alkohole aus dem Reaktionsmedium entfernt oder im wesentlichen entfernt werden und das so gebildete Gel getrocknet und gebrannt wird. 30

2. Verfahren nach Anspruch 1, dadurch gekennzeichnet, daß das Tetraalkylammoniumhydroxid Tetrapropyl-, Tetraisopropyl- oder Tetrabutylammoniumhydroxid ist, das Aluminiumtrialkoxid Aluminiumtripropoxid und -triisopropoxid ist und das 35

Tetraalkylorthosilicat Tetraethylorthosilicat ist. 40

3. Verfahren nach Anspruch 1 oder 2, dadurch gekennzeichnet, daß eine wäßrige Lösung, die Tetraalkylammoniumhydroxid und Aluminiumtrialkoxid enthält, hergestellt wird und zu dieser wäßrigen Lösung Tetraalkylorthosilicat zugegeben wird, wobei bei einer niedrigeren Temperatur als die Hydrolysetemperatur und mit solchen Mengen der Ausgangsreaktanten gearbeitet wird, die den Molverhältnissen SiO₂:Al₂O₃ im Bereich von 30:1 bis 500:1, Tetraalkylammoniumhydroxid:SiO₂ im Bereich von 0,05:1 bis 0,2:1 und H₂O:SiO₂ im Bereich von 5:1 bis 40:1 entsprechen, und die Hydrolyse und Gelbildung durch Erhitzen auf eine Temperatur im Bereich von etwa >65°C bis etwa 110°C in einem Autoklaven unter dem autogenen Druck des Systems oder unter Atmosphärendruck in einem mit Rückflußkühler ausgestatteten Reaktor bewirkt wird. 45

4. Verfahren nach Anspruch 1 bis 3, dadurch gekennzeichnet, daß ein freier Alkohol, vorzugsweise Ethanol, dem Reaktionsgemisch bis zu einem Maximalwert des Molverhältnisses von zugesetztem Alkohol:SiO₂ von 8:1 zugegeben wird. 50

5. Verfahren nach Ansprüchen 1 bis 4, dadurch gekennzeichnet, daß die Dauer der Hydrolyse/Gelbildung im Bereich von 10 Minuten bis 3 Stunden und vorzugsweise in der Größenordnung von 1 bis 2 Stunden liegt. 55

6. Verfahren nach Ansprüchen 1 bis 5, dadurch gekennzeichnet, daß das Gel der Alterung unterworfen wird, indem das Hydrolyse/Gelbildungs-Gemisch in Gegenwart des Alkohols bei Raumtemperatur während einer Dauer in der Größenordnung von 1 bis 24 Stunden gehalten wird. 60

7. Verfahren nach einem der vorhergehenden Patentansprüche, dadurch gekennzeichnet, daß der Alkohol von dem Gel entfernt wird und das letztere unter Vakuum getrocknet wird, und daß das getrocknete Gel dann in einer oxidierenden Atmosphäre bei einer Temperatur im Bereich von 500 bis 700°C während einer Dauer von 4 bis 20 Stunden, vorzugsweise bei 500 bis 600°C während 6 bis 10 Stunden gebrannt wird. 65

8. Verwendung des Siliciumdioxid-Aluminiumoxid-Gels, wie es nach Ansprüchen 1 bis 7 erhalten wird, in Verfahren zur Kohlenwasserstoffumwandlung. 70

9. Verwendung nach Anspruch 8, wobei das Kohlenwasserstoffumwandlungsverfahren die Propylen-Oligomerisation ist. 75

Revendications

1. Procédé de préparation d'un gel de silico-alumine micro/mésoporeux amorphe ayant une taille de pores contrôlée, une surface spécifique au moins égale à $500 \text{ m}^2/\text{g}$ et un rapport molaire $\text{SiO}_2/\text{Al}_2\text{O}_3$ au moins égal à 30:1, lequel procédé utilise comme substances de départ un hydroxyde de tétraalkylammonium, un composé d'aluminium capable de donner par hydrolyse du Al_2O_3 (c'est-à-dire hydrolysable en Al_2O_3), et un composé de silicium capable de donner par hydrolyse du SiO_2 (c'est-à-dire hydrolysable en SiO_2), dans lequel

- ledit hydroxyde de tétra-alkylammonium est un hydroxyde de tétra-(alkyle en C_{2-5})-ammonium, ledit composé d'aluminium hydrolysable est un tri-(alkylate en C_{2-4}) d'aluminium et ledit composé de silicium hydrolysable est un tétra-(alkyle en C_{1-5})-orthosilicate, et lesdits réactifs sont soumis à une hydrolyse et une gélification à une température supérieure ou égale à la température d'ébullition à pression atmosphérique de n'importe lequel des alcools qui se forment en tant que produits secondaires lors de ladite réaction d'hydrolyse, sans que lesdits alcools soient éliminés, ou éliminés de manière notable, du milieu réactionnel, et
- le gel ainsi produit est séché et cuit.

2. Procédé conforme à la revendication 1, caractérisé en ce que ledit hydroxyde de tétra-alkylammonium est l'hydroxyde de tétrapropylammonium, de tétrisopropylammonium ou de tétrabutylammonium, ledit trialkylate d'aluminium est du tripropylate d'aluminium ou du tri-isopropylate d'aluminium et ledit tétra-alkylorthosilicate est le tétraéthylorthosilicate.

3. Procédé conforme à la revendication 1 ou 2, caractérisé en ce que l'on prépare une solution aqueuse contenant de l'hydroxyde de téraalkylammonium et du trialkylate d'aluminium et en ce que l'on ajoute à ladite solution aqueuse du tétra-alkylorthosilicate en opérant à une température inférieure à la température d'hydrolyse et avec des quantités de réactifs de départ telles que le rapport molaire $\text{SiO}_2:\text{Al}_2\text{O}_3$ soit compris dans l'intervalle allant de 30:1 à 500:1, le rapport molaire hydroxyde de téra-alkylammonium: SiO_2 soit compris dans l'intervalle allant de 0,05:1 à 0,2:1 et le rapport molaire $\text{H}_2\text{O}:\text{SiO}_2$ soit compris dans l'intervalle allant de 5:1 à 40:1, et en ce que l'on effectue l'hydrolyse et la gélification par chauffage à une température comprise dans l'intervalle allant d'environ $>65^\circ\text{C}$ à environ 110°C , dans un autoclave à la pression autogène du système, ou à pression atmosphérique à l'intérieur d'un réacteur muni d'un condensateur à reflux.

4. Procédé conforme à la revendication 1 à 3, caractérisé en ce que l'on ajoute au mélange réactionnel l'alcool libre, et de préférence l'éthanol, jusqu'à une valeur maximale du rapport molaire alcool ajouté: SiO_2 égale à 8:1.

5. Procédé conforme aux revendications 1 à 4, caractérisé en ce que la durée d'hydrolyse/gélification est comprise dans l'intervalle allant de 10 minutes à 3 heures, et est de préférence de l'ordre d'une à deux heures.

6. Procédé conforme aux revendications 1 à 5, caractérisé en ce que l'on soumet le gel à un vieillissement en maintenant le mélange d'hydrolyse/gélification en présence de l'alcool et à la température ambiante pendant une durée de l'ordre de 1 à 24 heures.

7. Procédé conforme à l'une quelconque des revendications précédentes, caractérisé en ce que l'on élimine l'alcool du gel, en ce que l'on sèche ce dernier en opérant sous vide et en ce que l'on cuit ensuite le gel séché dans une atmosphère oxydante à une température comprise dans l'intervalle allant de 500 à 700°C pendant une durée comprise entre 4 et 20 heures, et de préférence à 500 - 600°C pendant 6 à 10 heures.

8. Utilisation du gel de silico-alumine obtenu conformément aux revendications 1 à 7 dans un procédé de conversion d'hydrocarbures.

9. Utilisation conforme à la revendication 8 dans lequel ledit procédé de conversion d'hydrocarbures est l'oligomérisation de propylène.

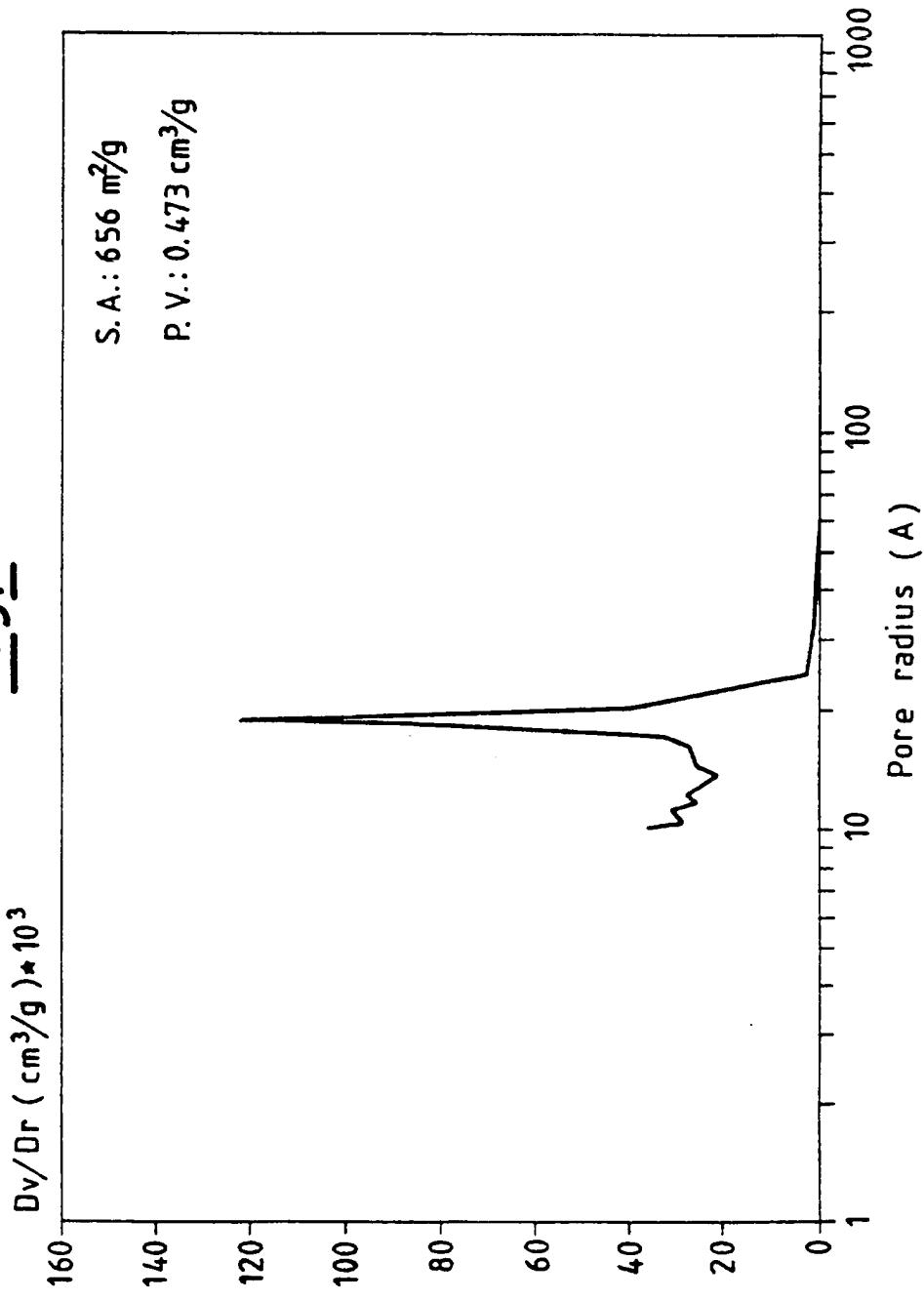
Fig.1

Fig.2

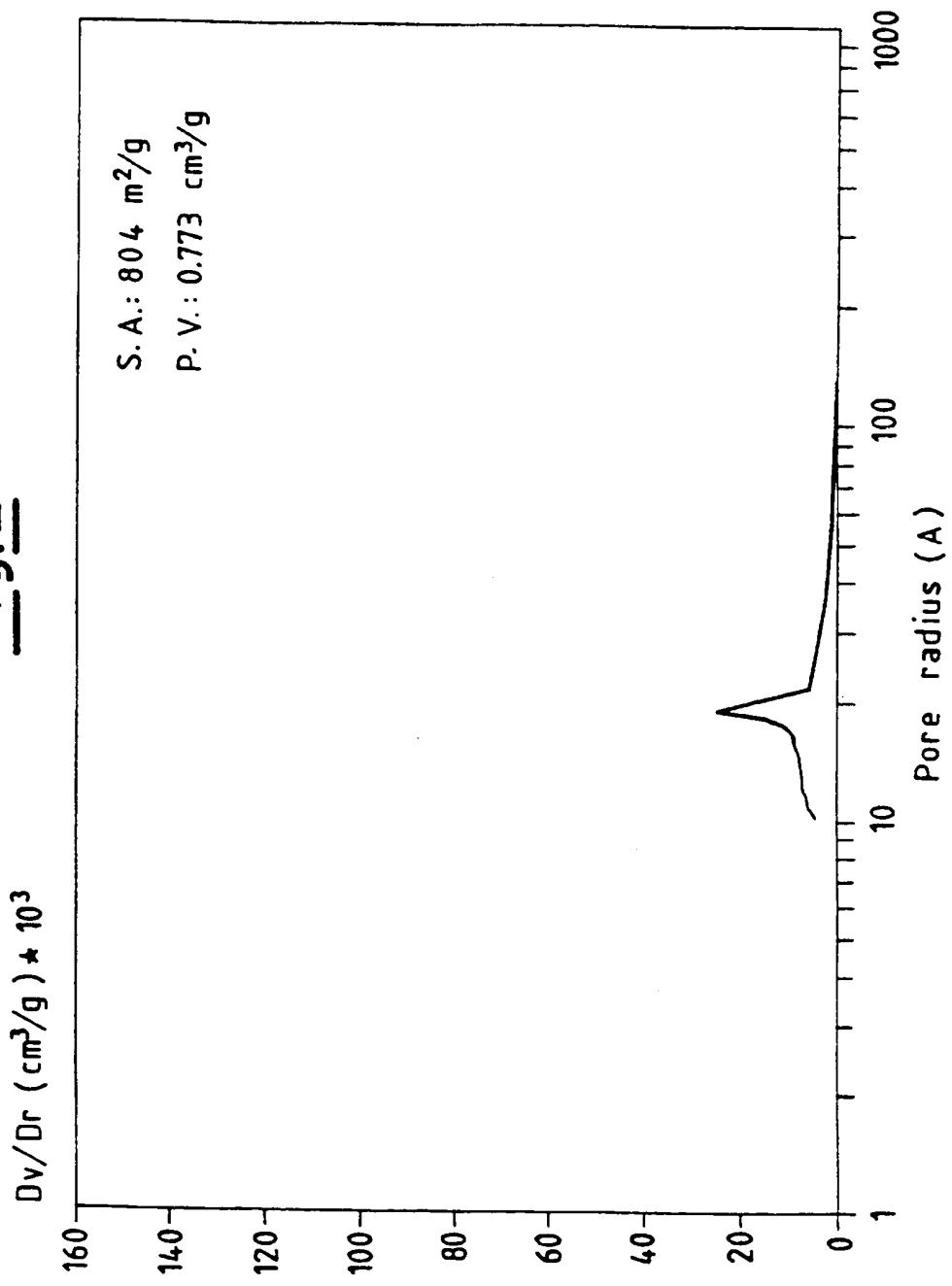


Fig. 3

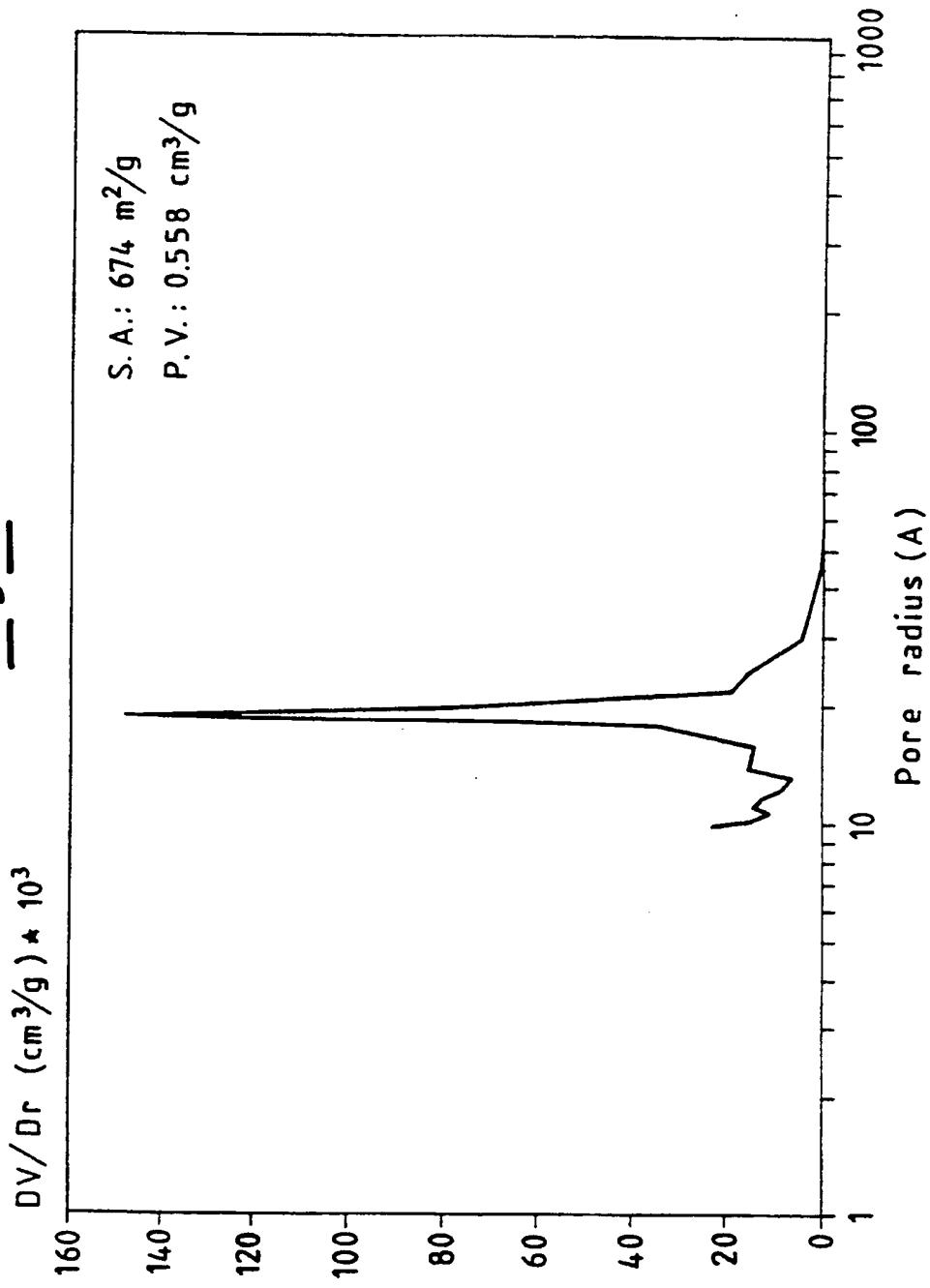


Fig.4



Fig.5

